SSLC -Chemistry -Class -17

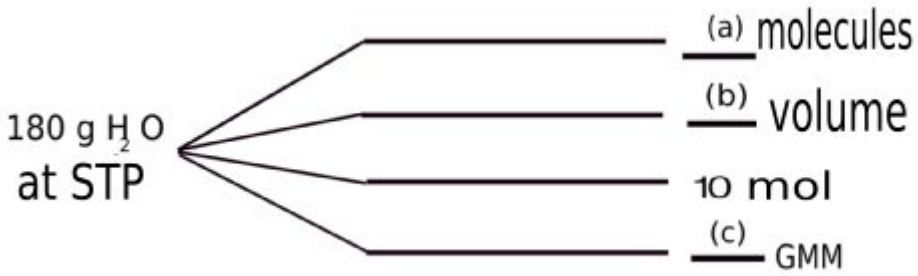
Unit 2 : Gas Laws and Mole Concept

Revision Questions

1) How many moles are there in 140g Nitrogen gas (N_2) ?

[10 , 5, 14. 140]

2. Find a ,b , c and d



3) 4 x 6.022 x10^{23} Chlorine molecules at STP are taken. Answer the following question (Atomic mass :Chlorine = 35.5)

a) What is its volume at STP ?

b) What is the mass of this compound?

c) How many moles are present in this sample ?

4) Match the following ?

A	B	C
54g H_2O	**3 X 6.022X10^{23}**	**1 Mole**
88g CO_2	**1 X 6.022X10^{23}**	**2 Moles**
17g NH_3	**2 X 6.022X10^{23}**	**3 Moles**

5) Choose the correct statement from those given below .

(Hint : Atomic mass : C - 12 , O - 16)

a) 6.022 x 10^{23} molecules are there in 22 g CO_2.

b) 1 GMM of CO_2 is 22 g .

c) Volume of 22 g CO_2 at STP is 11.2 L.

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