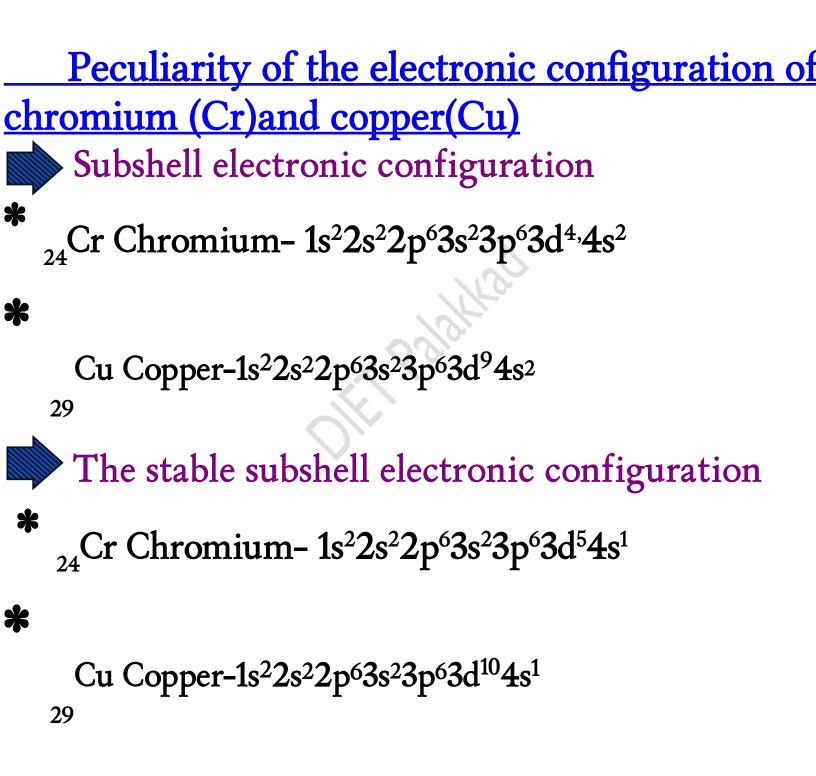


INTER BELL 2.0 / HS CHEMISTRY 10TH STD

<u>07-07-2021</u> WEDNESDAY



For the subshell electronic configurations of chromium and copper, the configurations with half filled d subshell or completely filled d subshell show greater stability.

Subshell electronic configuration and blocks

Based on subshell electronic configuration, elements

are classified into four blocks s,p,d,f

Sh	100	k															p b	loc	k	12
Н	2							dl	blog	ck					13	3 14	15	16	Vi	THE
Li	Be]													в	C	N	0	F	Ne
Na	Mg	1		3	4	5	6	7	8	9	10	11	12		A	I Si	Р	s	CI	Ar
к	Ca			Sc	Ti	v	Cr	Mn	Fe	Co	Ni	Cu	Zn		G	a Ge	As	Se	Br	Kr
Rb	Sr	1		Y	Zr	Nb	Mo	Te	Ru	Rh	Pd	Ag	Cd		In	n Sn	Sb	Те	T	Xe
Cs	Ba	1	[La	Hf	Та	w	Re	Os	lr	Pt	Au	Hg		т	1 Pb	Bi	Po	At	Rn
Fr	Ra	1		Ac	Rf	Db	Sg	Bh	Hs	Mt	Ds	Rg	Cn		N	h Fl	Mc	Lv	Ts	Og
							8	f bl	ocl	ĸ										
			Ce	Рт	Nd	Pm	Sm	Eu	Gd	Tb	Dy	Но	Er	Tm	Yb	Lu				
		You	ube)	w PR	pullu	e.Non	iišvi	að	Cm	Bk	Cf	E BA	<u>.</u>	****	v.filist	belEki	te.LeO	13,0	арп	n 🚚

INTERBELL WORKSHEET FOR FIRST BELL CLASSES

Period and Group can be found out on the

basis of subshell electronic configuration

	Configuration	which last electron is added	of outermost shell	electrons in outermost sub shell	Group No.	VICTION No.	
зLi	15 ² 25 ¹	s	2	1	1	2	
11Na	1s ² 2s ² 2p ⁶ 3s ¹	s	3	1	1	3	
20 ^{Ca}	1s ² 2s ² 2p ⁶ 3s ² 3p ⁶ 4s ²	s	4	2	2	4	

*For s block elements

Group number- The number of electrons in outermost s

subshell

Period number-Number of outermost shell

INTERBELL WORKSHEET FOR FIRST BELL CLASSES

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	Configuration	which last electron is added	Shell no. of outermost shell	No. of electrons in outermost sub shell	Group No.	Victoriod No. 2	
зLi	1s ² 2s ¹	S	2	1	1		
11Na	1s ² 2s ² 2p ⁶ 3s ¹	s	3	1	1	3	
20 Ca	1s ² 2s ² 2p ⁶ 3s ² 3p ⁶ 4s ²	s	4	2	2	4	

*For p block elements

- Group number-12+ number of electrons in outermost p
- subshell
- Period number-Number of outermost shell

ASSIGNMENT

- The atomic number of an element Y is 17.
- 1)Write the subshell electronic configuration of the element.
- 2)Find the period,group and block .