17/12/2020
THURSDAY

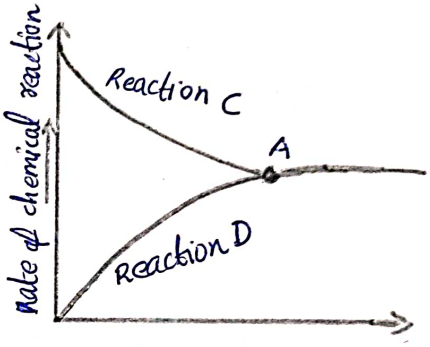
CHEMISTRY

STD - X
class - 28

Assignment

1. The graph for the reaction $N_2(g) + 3H_2(g) \rightleftharpoons 2NH_3(g) + Heat$ is given below.

   - Identify and write the reactions C and D.



Ans) Reaction C - Forward reaction
Reaction D - Backward reaction

2. $2NO(g) + O_2(g) \rightleftharpoons 2NO_2(g) + Heat$.

   In this reaction how do the following changes influence the rate of forward reaction?

Ans) a. When the concentration of oxygen is increased the rate of forward reaction increases. Hence more product is formed.

b. When increasing the concentration of NO, forward reaction increases.

c. When increasing the concentration of NO_2, backward reaction increases and forward reaction decreases.

d. When decreasing the concentration of NO_2, forward reaction increases.