



Online Class Supporting Materials

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CLASS: 10

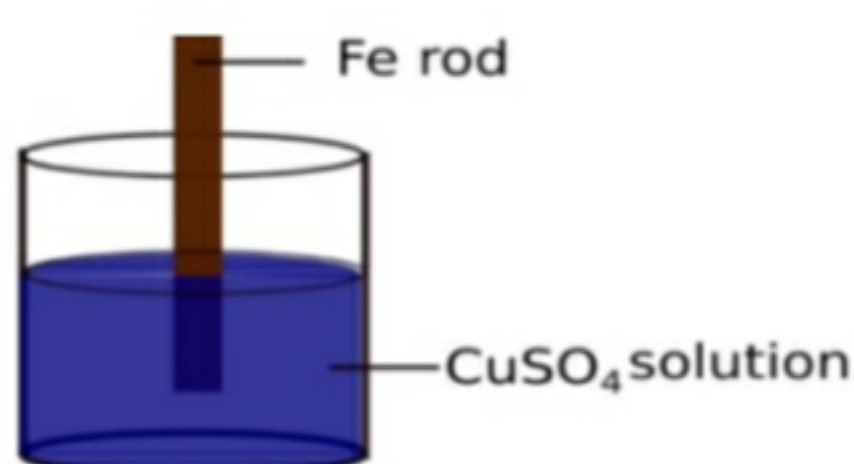
CHEMISTRY - 5

UNIT : REACTIVITY SERIES AND ELECTROCHEMISTRY

TOPIC:

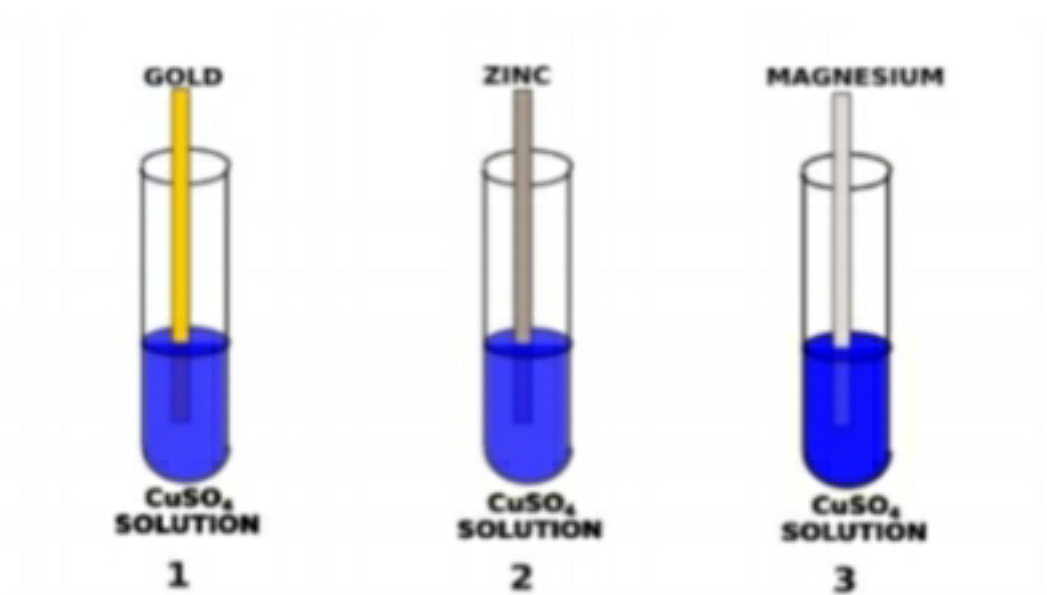
Reactivity series and displacement reactions.

1. Observe the picture.



- What are the changes that can be observed with the iron rod and the colour of the copper sulphate solution?
- Find the metal which was oxidised and the metal ion which was reduced. Write the equations of oxidation and reduction.
- Name the phenomenon which take place within the beaker.
- What will be the change if Silver rod is used instead of iron rod. What is the reason?
- Arrange the following metals in the increasing order of their reactivity.
i) Copper ii) Iron iii) Silver

2. Observe the picture.



- In which test tube displacement reactions can be observed?
- Write the redox equation for each reaction.

TOPIC: GALVANIC CELL

- 1) Find the energy change that takes place in Galvanic cell?
 - a) Electrical energy changed to chemical energy
 - b) Chemical energy changed to electrical energy
 - c) Electrical energy changed to light energy

- 2) Select the materials needed for making Zn-Cu Galvanic cell

(Test tubes, 2 beakers, Zn, Cu, Mg metal rods, ZnSO_4 , MgSO_4 , CuSO_4 salt solution, copper wire, salt bridge, voltmeter)

- 3) Write down the chemical reaction that takes place in the anode of Cu-Ag Galvanic cell



4) The following materials are given.

Rods of Zn, Cu, Ag, solutions of ZnSO_4

CuSO_4 , AgNO_3

(Hint: reactivity order $\text{Zn} > \text{Cu} > \text{Ag}$)

a) Which metal rod is only used as anode

b) Which metal rod is only used as cathode

c) In which electrode reduction take place in the Cu-Ag galvanic cell

d) Write down the chemical reaction takes place in the anode of Cu-Ag cell.

5) Draw a Zn-Cu galvanic cell and mention the important parts such as anode, cathode, direction of electron flow.