CWX2 (3)

CHAPTER 2 MODULE 3
CHEMISTRY
STANDARD X

1.   **128 gm O_2 = ……..GMM   (Atomic mass of oxygen=16)**

2.   **Find the molecular mass of the following compounds (Atomic mass of the elements are Na - 23, O- 16, H- 1, Ca- 40, C- 12, N- 14)**

a)   **NH_3**

b)   **$CaCO_3$**

c)   **NaOH**

3.   **Find the pair (Atomic mass O-16, H-1, Ca-40, C-12)**

36 g H_2O

3 GMM

132 gm CO_2

3.011×10^{23} molecules

50 gm Ca CO_3

2 Mole

4.   **Complete the Table**

1 GMM = 1 Mole = 6.022×10^{23} molecules

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Element/ Compound	Gram Molecular Mass	Mass in gram	No. of moles	No.of molecules
Hydrogen	2	6	3	$3 \times 6.022 \times 10^{23}$
Carbon di Oxide	44	-----	2	--------
Sulphuric acid	---	490	5	$5 \times 6.022 \times 10^{23}$
Calcium Carbonate	---	500	------	$5 \times 6.022 \times 10^{23}$

5. **Volume of 1 mole of any gas at STP = 22.4 L**

Gas at STP	Gram Molecular Mass	Mass in gram	Moles	Volume at STP
CO_2	44	220	5	$5 \times 22.4L$
H_2	2	------	6	---------L
NH_3	----	170	10	--------L
N_2	-----	112	-----	$4 \times 22.4L$

6   **Complete the DIAGRAM**