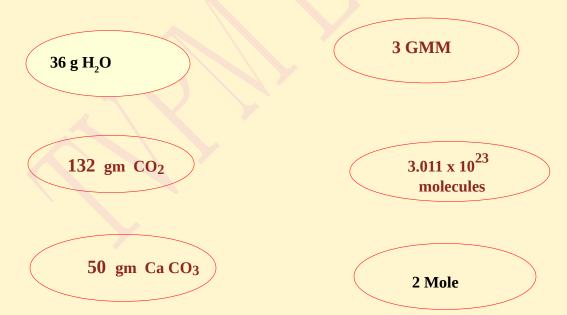
## THIRUVANANTHAPURAM EDUCATIONAL DISTRICT



## CHAPTER 2 MODULE 3 CHEMISTRY STANDARD X

- 1. 128 gm  $O_2 = \dots GMM$  (Atomic mass of oxygen=16)
- 2. Find the molecular mass of the following compounds (Atomic mass of the elements are Na 23, O- 16, H- 1, Ca- 40, C- 12, N- 14)
- a) NH<sub>3</sub>
- b) CaCO<sub>3</sub>
- c) NaOH
- **3.** Find the pair (Atomic mass O-16, H-1, Ca-40, C-12)



4. Complete the Table

1 GMM = 1 Mole =  $6.022x10^{23}$  molecules

Element/ Compound	Gram Molecular Mass	Mass in gram	No. of moles	No.of molecules
Hydrogen	2	6	3	3x6.022x10 <sup>23</sup>
Carbon di Oxide	44		2	
Sulphuric acid		490	5	5x6.022x10 <sup>23</sup>
Calcium Carbonate		500		5x6.022x10 <sup>23</sup>

## 5. Volume of 1 mole of any gas at STP = 22.4 L

Gas at STP	Gram Molecular Mass	Mass in gram	Moles	Volume at STP
CO <sub>2</sub>	44	220	5	5x22.4L
$H_2$	2		6	L
$NH_3$		170	10	L
$N_2$		112		4x22.4L

## 6 Complete the DIAGRAM

