## FIRST TERM MODEL EXAMINATION 2020-21

STD:X
CHEMISTRY
Time:1 1/2 HRS
Marks:40

Questions 1 to 5 carries 1 marks. 6 to 10 carries 2 marks. 11 to 15 carries 3 marks. 16 to 20 carries 4 marks.(from each section write any 4 answers)

## Section-A

1.Among the following elements ,which one show the +2 oxidation state ( $\mathrm{Fe}, \mathrm{Ag}, \mathrm{K}$ )
2.The most electronegative element is
(Au, F, Na)
3.One Avogadro number is equal to
$\left(\begin{array}{ll}6.02 \times 10^{21} & 6.02 \times 10^{22} \\ 6,02 \times 10^{23}\end{array}\right)$
4. What is the GMM of 64 gm of oxygen
(2, 32, 64)
5.Reaction between Zn and $\mathrm{CuSO}_{4}$, oxidation will happen to ---$\left(\mathrm{Cu}, \mathrm{Zn}, \mathrm{SO}_{4}\right)$
Section-B
6.fresh magnesium ribbon losing its luster
when kept exposed in the air for some days. Why?
7. Calculate the number of molecules present in $\mathbf{9 0} \mathrm{gm}$ of water?
8.Write any 2 examples of charle's law.
9. Can you explain why d-block elements shows different oxidation states?
10.What is the family name of elements belonging to the 17 th group? What is their common valency?

## Section-C

11. Calculate the mass of 112 L CO 2 gas kept at STP (molecular mass =44). How many molecules of $\mathrm{CO}_{2}$ are present in it?
12. The size of the air bubbles rising from the bottom of an aquarium increases. Can you explain the reason?Write any 2 practical applications of this law.
13. What happened when a ' Mg ' rod is dipped in ${ }^{\text {' } \mathrm{CuSO}^{4} \text { ' }}$ solution. Write down the chemical equation?
14. Write down the electronic configuration of chromium?
15. The electronic configuration of an element is $1 s^{2} 2 s^{2} 2 p^{5}$. Find the block and period of the element?

## Section-D

16.The Molecular mass of ammonia is 17.
a) How much is the GMM of ammonia?
b) Find out the number of moles of molecules present in 170 g of ammonia?
17.

| Gas | Volume (L) | No. of molecules |
| :--- | :---: | :---: |
| Nitrogen | 10 L | x |
| Oxygen | 5 L |  |
| Ammonia | 10 L | --- |
| Carbon di oxide | --- | 2 x |

a)Complete the table.
b) Which gas law is applicable here?
18. Draw a galvanic cell made of Zn and Cu . Write down the anode and cathode reactions?
19. Write the electronic configuration of $\mathrm{Zn}^{+2}$. Find the group and period of Zn ?
20. Write any 3 properties of d- block elements?

