1.

Fill in the blanks
1 Dozen = 12 No.s
1mole   = 6.022 X 10^{23} .No.s

2.
Complete the table

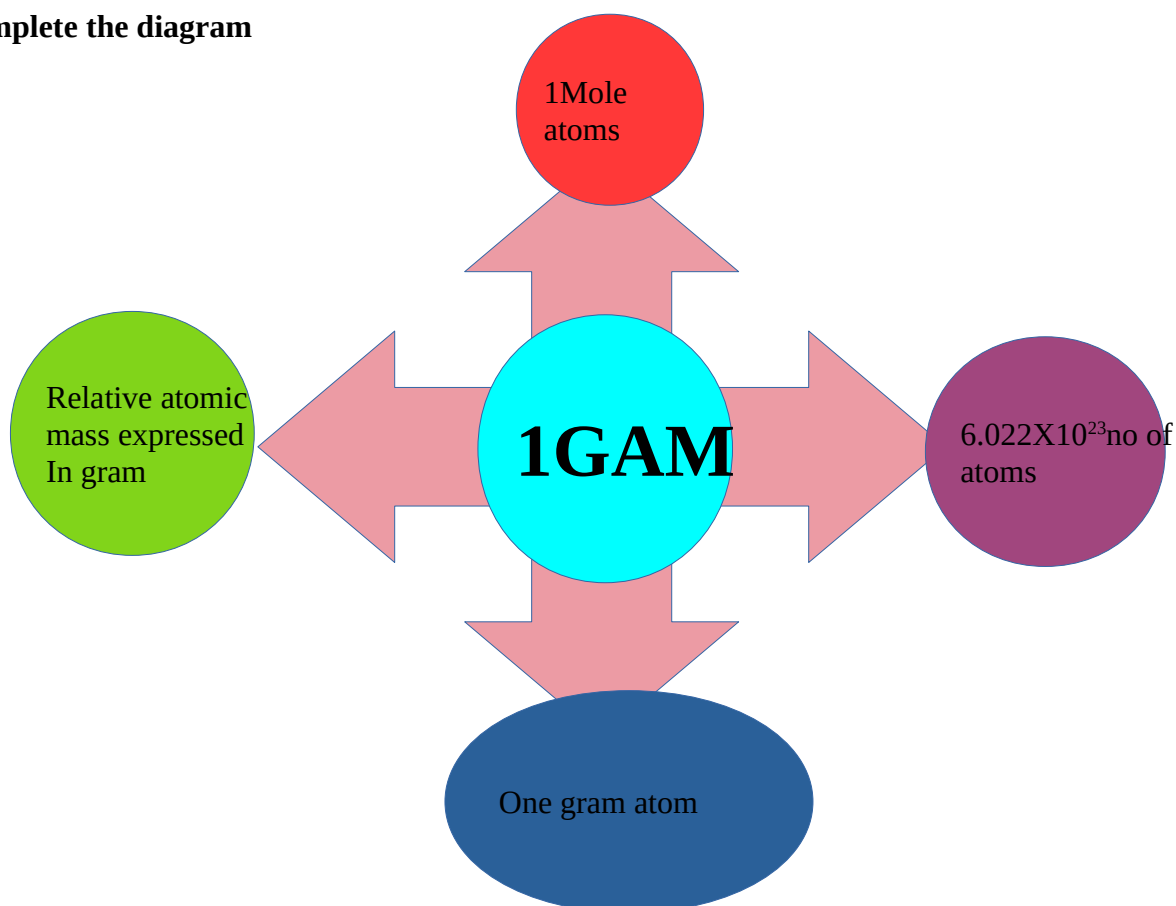| Element	Relative atomic mass	GAM (Relative atomic mass in gram)	No.s of atoms in 1GAM
Carbon	12	12g	6.022 X 10^{23}
Neon	20	20g	6.022 X 10^{23}
Calcium	40	40g	6.022 X 10^{23}
Sulphur	32	32g	6.022 X 10^{23}

3.

Find out the pairs having equal no.s of atoms
➢ 10 g Hydrogen
➢ 140 g Nitrogen
➢ 230g Sodium

4.
Complete the diagram

5.

    Find out no. of atoms present in following samples

(a) 240g of carbon - $20 \times 6.022 \times 10^{23}$

(b) 460g Sodium   - $20 \times 6.022 \times 10^{23}$