

CBSE Class 10 Science
Ch-5 Periodic Classification of elements

More Questions

1. State the modern periodic law
2. which of the two elements
3. How does the atomic size vary in going from
 - A. left to right in a period
 - B. Top to Bottom in a group
4. An element has atomic number 13. In which group and period it should be placed?
5. How many periods and groups are there in the long form of P.T?
6. Why does the size of the atoms progressively become smaller when we move from sodium (Na) to chlorine (Cl) in the third period of the periodic table?
 - A. Give symbols for
 - B. A metal of group 2.
 - C. A metal of group 13.
 - D. Two of group 13.
 - E. name of the element
 - F. Relative size with respect to other members of its group.
7. Explain Why-
 1. All the elements of a group have similar chemical properties.
 2. All the elements in a period have different chemical properties.
8. The atomic number of an element X is 17. Predict-
 - A. Its valency.
 - B. Nature of the elements.
 - C. Name of the element is metal or non-metal.
 - D. Name of the element.
 - E. Relative size with respect to other members of its group.
9. The three elements predicted by Mendeleev from the gaps in his periodic table were known as eka-boron, eka-aluminium, eka-silicon. What names were given to these elements when they were discovered later on?
10. The atomic numbers of Nitrogen, Oxygen and fluorine are 7, 8 and 9 respectively. Write

the electronic configuration of each element and answer the following:

- a. Out of N, O and F which is most electronegative and which one is least electronegative?
- b. What is the number of valence electrons of F?
- c. What is the valency of each of N, O and F?